

Ionic Dot Diagrams

- An Ion is an atom with a positive or negative charge
- The size and type of charge depend on the number of valence electrons
- Atoms in group 1(Na), 2(Ca) and 13(B) LOSE electrons to become positive
- Their 'normal' dot diagrams look like this:

- Their Ionic dot diagram looks like this:

- **Note that instead of 'dots' you write the type of charge (+) and the size (number) next to the atomic symbol**
- Atoms in group 15(N), 16(O) and 17(Cl) GAIN electrons to become negative
- Their 'normal' dot diagrams look like this:

- Their Ionic dot diagram looks like this:

- **Note that instead of 'dots' you write the type of charge (-) and the size (number) next to the atomic symbol**
- When a NON-METAL bonds with a METAL, they will make IONIC bonds
- The size of the charge determines how many bonds they can form