Molecular Atomic Inventory Practice

1) H_2O

H-_____

O-_____

2) CO_2

C-_____

O-_____

3) NH_3

N-_____

H-_____

4) H_2O_2

H-_____

O-_____

5) C_2H_6

C-_____

H-_____

6) HCN

H-_____

C-_____

N-_____

7) NaCl

Na-_____

Cl-_____

8) PO_3

P-_____

O-_____

9) CrO_2

Cr-_____

O-_____

10) MnO_4

Mn-_____

O-_____

11) ClO_4

Cl-_____

O-_____

12) AlO_2

Al-_____

O-_____

13) HPO_4

H-_____

P-_____

O-_____

14) SiO_3

Si-_____

O-_____

15) AsO_3

As-_____

O-_____

16) IF

I-_____

F-_____

17) $C_6H_{12}O_6$

C-_____

H-_____

O-_____

18) H_2SO_4

H-_____

S-_____

O-_____

19) $Mg(OH)_2$

Mg-_____

O-_____

H-_____

20) CH_2OH

H-_____

O-_____

H-_____

21) H_3PO_4

H-_____

P-_____

O-_____

22) $CuCl_2$

H-_____

O-_____

23) $MgSO_4$

H-_____

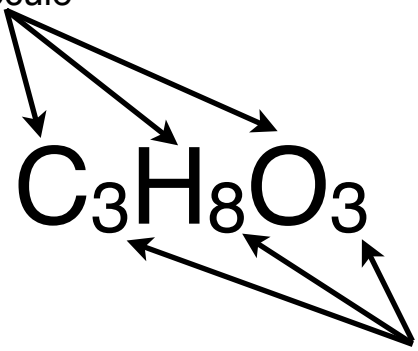
S-_____

O-_____

Atomic Inventories

Scientists use chemical formulas to show how many atoms are found in a molecule

The letters tell you what atoms make up the molecule



The small numbers tell you how many of each atom there are in the molecule

An atomic inventory is a way of counting the atoms in a molecule

$C_3H_8O_3$

C-_____

H-_____

O-_____

In each blank you write the number of atoms that are in the molecule.
For example:

$C_3H_8O_3$

C- 3

H- 8

O- 3