Solutions and Solubility Notes

- **Solute-**the solid that is added to a liquid to make a mixture
- **Solvent-**the liquid that the solid is added to make a mixture
- **Suspension-**a mixture where the solute can be seen and easily separated from the solvent
- **Solution-**a mixture where the solute cannot be easily seen and separated from the solvent
- **Solubility-**how much solute can be mixed into an amount of solvent
- **Concentration-**a measure of how much solute is in a mixture

- Solubility is dependent on the temperature of the solution
    - Lower temperatures decrease the amount of solute a solution can hold
    - Higher temperatures increase the amount of solute a solution can hold

- Concentration is a way of describing the amount of solute that is in a solution
- Scientist use brackets-[ ] to abbreviate concentration
    - So the concentration of Sodium in a solution can be written as [Na]
    - [Cl] would mean ‘the concentration of Chlorine’
- Scientists change concentration by adding more solute to the solution
    - **Dilute Solution**-small amount of solute
    - **Concentrated Solution-**large amount of solute

- When a solution cannot hold anymore solute, we call it a **Saturated Solution**
- When a solution can hold more solute, we call it an **Unsaturated Solution**